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CHAPTER 9 REVIEW Stoichiometry

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percentage yield. 2. 6.0 mol of N_2 are mixed with 12.0 mol of H_2 according ...

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Review Vocabulary reactant: the starting substance in a chemical reaction
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ANSWER Answer the following questions in the space provided. 1. _____ The coefficients in a chemical equation represent the (a) masses in grams of all reactants and products. (b) relative number of moles of reactants and products. (c) number of atoms of each element in each compound in a reaction.

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Stoichiometry. SECTION 2. PROBLEMS

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Write the answer on the line to the left.
Show all your work in the space
provided.

1. The following equation
represents a laboratory preparation for
oxygen gas: $2\text{KClO}_3(\text{s}) \rightarrow 2\text{KCl}(\text{s}) + 3\text{O}_2(\text{g})$
How many moles of O_2 form if
3.0 mol of KClO_3 are totally consumed?

2. Given the following equation: $\text{H}_2(\text{g}) + \text{F}_2(\text{g}) \rightarrow 2\text{HF}(\text{g})$

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CHAPTER 9 REVIEW

CHAPTER 9 REVIEW Stoichiometry

SECTION 3 PROBLEMS Write the answer on the line to the left. Show all your work in the space provided.

1. 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield.

2. 6.0 mol of N₂ are

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mixed with 12.0 mol of H₂ according to the following equation: N

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1 mol Al₂O₃ or 1 mol Al₂O₃ ___
101.96 g Al₂O₃ 26.98 g Al _ 1 mol Al
or _ 1 mol Al 26.98 g Al 32.00 g O₂ _ 1
mol O₂ or 1 mol O₂ _ 32.00 g O₂ To

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find the number of grams of aluminum equivalent to 26.0 mol of aluminum, the calculation would be as follows. 26.0 mol Al \times 26.98 g Al / 1 mol Al = 701 g Al
Stoichiometry 291 SECTION 1 ...

**CorrectionKey=NL-A DO NOT
EDIT--Changes must be made ...**

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1. Write the definition of reaction stoichiometry in your own words.

Introduction to Stoichiometry SECTION

9.1 amount of given substance (mol)

convert into amount of unknown

substance (mol) Ratios of substances in

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chemical reactions can be used as conversion factors. Reaction stoichiometry problems can be approached by looking

SECTION 9.1 Introduction to Stoichiometry

SECTION 2 continued Date Class _____
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100. of lithium carbonate. The oxygen gas produced in part a has density of 1.43 g/L calculate the volume of this gas. 76
STOICHIOMETRY MODERN CHEMISTRY a. —. 81 g 6. A car air bag requires 70. L of nitrogen gas ...

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