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the following questions in the space

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provided. 1. Given the following equation: $C_3H_4(g) + xO_2(g) \rightarrow 3CO_2(g) + 2H_2O(g)$ a. What is the value of the coefficient x in this equation? 40.07 g/mol b. What is the molar mass of C_3H_4 ? 2 mol O_2 :1 mol H_2O c. What is the mole ratio ...

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CHAPTER 9 REVIEW Stoichiometry

SECTION 3 PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1. _____ The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield. 2. 6.0 mol of N₂ are

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mixed with 12.0 mol of H

CHAPTER 9 REVIEW Stoichiometry

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Section 3

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9 REVIEW Stoichiometry SECTION 3
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space provided 1 88% The actual yield
of a reaction is 22 g and the theoretical
yield is 25 g Calculate the percentage

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yield 2 60 mol of N_2 are mixed with 120 mol of H

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