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Chapter 1 Significant Figures  
Worksheet

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$1.423 \times 4.2 = 6.0$  since 1.423 has 4 significant figures and 4.2 only has two significant figures, the final answer must also have 2 significant figures.  $234.67 - 43.5 = 191.2$  since 43.5 has one decimal place and 234.67 has two decimal places, the final answer must have just one decimal place.

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#### **Significant Figures | Introduction to Chemistry**

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Significant Figures Worksheet  $10^3$ ),  
with three significant figures. A number  
less than one, such as 0.000416, can be  
written in scientific notation as  $(4.16$



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#### **Chemistry Chapter 1 Significant Figures Worksheet**

The measurement 140 can be written as  $(1.4 \times 10^2)$ , with two significant figures in the coefficient or as  $(1.40 \times 10^3)$ , with three significant figures. A number less than one, such as 0.000416, can be written in scientific

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notation as  $(4.16 \times 10^{-4})$ , which has 3 significant figures.

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By rule 1, all nonzero digits are significant, so this measurement has three significant figures. By rule 4, the first three zeros are not significant, but by rule 2 the zero between the sixes is; therefore, this number has four significant figures.

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number of significant figures for the three number is \_\_\_\_\_. (a) 3,3 and 4 respectively (b) 3,4 and 4 respectively (c) 3,4 and 5 ...

### **Class 11 Important Questions for Chemistry chapter -1**

So, by these rules, the population figure of the United States has only three

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significant figures: the 3, the 6, and the zero between them. The remaining six zeros simply put the 306 in the millions position. (See Figure 1.7 "Significant Figures" for another example.)

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## **SIGNIFICANT FIGURES || CLASS 11**

### **Chapter 02|| Units and ...**

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<https://youtu.be/7hhW3sbHkf0> We will talk about Significant Figures and what is their importance in Chemistry...

### **Significant Figures - CBSE Class 11 Chemistry - YouTube**

The 1 and 100 in the conversion factors do not affect the determination of significant figures because they are

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exact numbers, defined by the centi-prefix. Test Yourself What is the volume of a block in cubic meters whose dimensions are  $2.1 \text{ cm} \times 34.0 \text{ cm} \times 118 \text{ cm}$ ?

### **Converting Units - Introductory Chemistry - 1st Canadian ...**

Chapter 1: Matter and Energy 25

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Significant Figures • The total number of digits in a measurement is called the number of significant figures. • When reading a scale, the value you record should use all of the digits you are sure of, plus one additional digit that you estimate. This last estimated digit is the last significant figure in your

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**Chapter 1 Matter and Energy -  
Angelo State University**

Chemistry Chemistry Indicate the number of significant figures in each of the following: a. This book contains more than 1000 pages. b. A mile is about 5300 ft. c. A liter is equivalent to 1.059 qt. d. The population of the United States is approaching  $3.1 \times 10^2$  million.

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#### **Indicate the number of significant figures in each of the ...**

The least number of decimal places is critical for finding significant figures after addition and subtraction functions, whereas the least number of significant figures within the starting numbers is important for multiplication and division

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functions. Recall that only measured numbers are used in the determination of significant figures.

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The number of significant figures present in give number has to be given.  
Concept introduction: Significant figures:  
The digits having a meaning in a given



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number are called as significant figures.  
Significant figures – calculation rules.  
Significant figure is a non-zero number  
(4443.2 has five significant figures)

**How many significant figures are in each of the following ...**

Figure 1.2: (Upper Panel) The Periodic Table of the Elements is an organized

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chart that contains all of the known chemical elements.(Lower Panel) To the left of the arrow is shown one atom of oxygen and two atoms of hydrogen.Each of these represent single elements. When they are combined on the righthand side, they form a single molecule of water ( $H_2O$ ).

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