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1000 g = 1 kg. 1000 mL = 1 L. Taking the 1.15 molal NH4OH in 1200 g solvent, we solve for the mass of solute. 1200g solvent \times 1 kg 1000g \times 1.15 mol solute kg solvent. = 1.38 mols solute. where the subscript indicates the digit past the last significant digit.

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CHAPTER 12 REVIEW Solutions

Chapter 12 Problem Set 3 1. From the following enthalpy changes, same C (s) + O2 (g) CO (g) H = 110.5 kJ same CO (g) + O2 (g) CO2 (g) H = 283.0 kJ calculate the value of H for the reaction C (s) + O2 (g) CO2 (g).

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2 Answers. the combustion of methane is an exothermic process. therefore the reactants must posses higher/lower potential energy than the reactant. 1 Answer. ASK A QUESTION. Remove ads. Upgrade to premium! There was an error saving. Please reload the page.

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238 Chemistry: Matter and Change • Chapter 12 Solutions Manual CHAPTER 12 SOLUTIONS MANUAL 7. Challenge Air is a mixture of gases. By percentage, it is roughly 78 percent nitrogen, 21 percent oxygen, and 1 percent argon. (There are trace amounts of many other gases in air.) If the atmospheric pressure is 760 mm Hg, what