

# Answers For Chemistry Module 13

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### **Answers For Chemistry Module 13**

In order to answer any question about the energy involved in a chemical reaction, we must first get  $\Delta H$ . Since Table 13.2 contains all needed info., we will use Hess's Law:  $\Delta H^\circ = (2 \text{ moles}) \times (-45.9 \text{ kJ/mole}) + (4 \text{ moles}) \times (-242 \text{ kJ/mole}) - (2 \text{ moles}) \times (33.2 \text{ kJ/mole})$ .  $\Delta H^\circ = -1,126 \text{ kJ}$ .

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In chemistry we measure matter using \_\_\_\_ units. This is an abbreviation for ... STEP 4: Now, use your answer from Step 3 as the new "HAVE" and repeat the process using the conversion factor 60 minutes = 1 hour ... and the lowest is -5.4% that range is 13.0%, which means that experiment was not precise Practice: Cheryl, Cynthia, Carmen, and ...

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### **Answer 1 Mark Page 9 of 13 Chemistry 30 Unit B Module 4 ...**

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